Heterogenous CO_2 and CH_4 content of glacial meltwater from the Greenland Ice Sheet and implications for subglacial carbon processes

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Abstract. Accelerated melting of the Greenland Ice Sheet has increased freshwater delivery to the Arctic Ocean and amplified the need to understand the impact of Greenland Ice Sheet meltwater on Arctic greenhouse gas budgets. We measured carbon dioxide (CO_2) and methane (CH_4) concentrations and $\delta^{13}C$ values and use geochemical models to evaluate subglacial CH_4 and CO_2 sources and sinks in water discharging from three subglacial outlets of the Greenland Ice Sheet in southwest (a sub-catchment of the Isunnguata Glacier; sub-Isunnguata, and the Russell Glacier) and southern Greenland (Kiattut Sermiat). Meltwater CH_4 concentrations vary by orders of magnitude between sites and are saturated with respect to atmospheric concentrations at Kiattut Sermiat. In contrast, meltwater in southwest sites is supersaturated, even though oxidation reduces CH_4 concentrations by up to 50% during periods of low discharge. CO_2 concentrations range from supersaturated at sub-Isunnguata to undersaturated at Kiattut Sermiat. CO_2 is consumed by mineral weathering throughout the melt season at all sites, however differences in the magnitude of subglacial CO_2 sources result in meltwaters that are either sources or sinks of atmospheric CO_2. At the sub-Isunnguata site, the predominant source of CO_2 is organic matter (OM) remineralization. At Russell and Kiattut Sermiat sites, multiple or heterogeneous subglacial CO_2 sources that maintain atmospheric CO_2 concentrations at Russell but not at Kiattut Sermiat where CO_2 is undersaturated. These results highlight a previously unrecognized degree of heterogeneity in greenhouse gas dynamics under the Greenland Ice Sheet. Future work should constrain the extent and controls of heterogeneity to improve our understanding of the impact of Greenland Ice Sheet melt on Arctic greenhouse gas budgets, as well as the role of continental ice sheets in greenhouse gas variations over glacial-interglacial timescales.

1 Introduction

Glaciers play an important role in global chemical cycles due to the production of fine-grained sediments that participate in carbonate and silicate mineral weathering reactions (Table 1), which are the principal sink of atmospheric CO_2 over geologic timescales (Berner et al., 1983; Walker et al., 1981). Variations in the weathering intensity of comminuted

sediments may contribute to glacial-interglacial atmospheric CO_2 variations as sediments are alternately covered by ice and exposed following ice retreat. However, the importance of CO_2 consumption by mineral weathering is poorly understood, including effects from the advance and retreat of continental ice sheets (Ludwig et al., 1999). Recent evaluations of carbon budgets in proglacial environments indicate that mineral weathering results in net sequestration of atmospheric CO_2, suggesting that proglacial systems are underrecognized as Arctic CO_2 sinks (St Pierre et al., 2019), however alternate processes could lead to the production of greenhouse gases in glacial systems. For instance, CH_4 production in anaerobic subglacial environments driven by the remineralization of organic matter (OM) contained in soils and forests covered during glacial margin fluctuations has been suggested as a potential carbon feedback to drive warming (Sharp et al., 1999; Wadham et al., 2008). Because the global warming potential of CH_4 is 25 times greater than CO_2, even limited subglacial methanogenesis has the potential to strongly impact the greenhouse gas composition of glacial meltwater. Combined inorganic and organic subglacial processes may therefore produce glacial meltwater that is a source or sink of greenhouse gas. While the net impact of these processes on modern carbon fluxes is poorly constrained, determining these impacts will improve modern carbon budgets as well as depictions of how fluxes may have evolved during the advance and retreat of continental ice sheets.

In subglacial environments where remineralization is limited by low OM availability, the major element solute load of glacial meltwater is typically dominated by products of mineral weathering reactions (Tranter, 2005). The extent of mineral weathering in subglacial environments depends in part on the availability of acids to drive reactions, namely sulfuric and carbonic acids (Table 1). Sulfuric acid is derived from the oxidation of reduced sulfur species, which largely occur as iron-sulfide minerals including pyrite (Tranter, 2005). Sulfide oxidation may occur abiotically; however, the kinetics of microbially mediated sulfide oxidation is several orders of magnitude faster and may lead to local depletion of oxygen given a sufficient supply of sulfide minerals (Sharp et al., 1999). In contrast, carbonic acid may be derived from multiple external or *in situ* sources of CO_2 to the system. The dominant external source is supraglacial meltwater that flows to the subglacial system through moulins following equilibration with atmospheric CO_2 (Fig. 1). Unlike proglacial environments where free exchange between water and the atmosphere may resupply CO_2 consumed by weathering, subglacial environments may be partially or fully isolated from the atmosphere, limiting further atmospheric CO_2 invasion and thus the extent of mineral weathering with carbonic acid. However, additional atmospheric CO_2 may be delivered in open portions of the subglacial environment though exchange in fractures or moulins along subglacial flow paths or in partially air-filled conduits, allowing a much greater magnitude of carbonic acid mineral weathering (Graly et al., 2017). CO_2 may also be derived from *in situ* sources, such as gaseous CO_2 contained in ice bubbles of basal ice, or fluid inclusions in rocks that release volatiles (including CO_2) following mechanical grinding (Macdonald et al., 2018). When OM is available, its remineralization also generates CO_2 (and potentially CH_4) along with nutrients, but low OM availability in many subglacial systems limits remineralization as a CO_2 source (Fig. 1).

The role of subglacial carbon processes may play an increasingly important role in modern Arctic carbon budgets as disproportionate warming increases glacial meltwater and sediment fluxes to the ocean, particularly from the Greenland Ice

Sheet. The is the last remaining ice sheet in the Northern hemisphere following collapse of other ice sheets since the Last Glacial Maximum (~20 ka). It has been losing mass at increasing rates that averaged 286±20 Gt/yr between 2010-2018, representing a six-fold increase since the 1980s (Mouginot et al., 2019). While mineral weathering significantly modifies the chemical composition of Greenland Ice Sheet subglacial discharge (e.g. Hindshaw et al., 2014; Deuerling et al., 2018; Urra et al., 2019) and should consume CO_2 similar to other glacial and proglacial environments, the recent identification of microbially driven reactions (including methanogenesis) in subglacial environments of the Greenland Ice Sheet indicates that organic processes may also play a role (Dieser et al., 2014; Lamarche-Gagnon et al., 2019; Musilova et al., 2017). The relative importance of subglacial greenhouse gas sinks (CO_2 consumption through mineral weathering) and sources (such as OM remineralization) determine the greenhouse gas composition of subglacial discharge, which may then serve as a source or a sink of atmospheric greenhouse gases. Constraining the relative impacts and variability of these processes underneath the Greenland Ice Sheet will provide important information regarding the current and future impact of Greenland Ice Sheet loss on Arctic carbon budgets, as well the role of continental ice sheets on carbon cycle feedbacks.

To evaluate the net impact of carbon processes on the greenhouse gas composition of subglacial discharge of the Greenland Ice Sheet, we compare water chemistry, dissolved CO_2 and CH_4 concentrations, and gas stable isotopic compositions between three subglacial discharge sites draining land-terminating glaciers of the Greenland Ice Sheet over the melt seasons of 2017 and 2018 (Fig. 2). We employ mass balance models utilizing the concentrations of major cations and anions to determine the magnitude of the impact on CO_2 concentrations from mineral weathering reactions (Table 1). These results are combined with measured gas concentrations and $\delta^{13}C$ to determine the relative importance of mineral weathering compared to OM remineralization on the CH_4 and CO_2 content of subglacial discharge. We also assess the temporal and spatial variability of these processes under the Greenland Ice Sheet to improve our understanding of carbon cycling in Greenland subglacial environments and the implications of Greenland Ice Sheet mass loss on Arctic carbon budgets.

2 Methods

2.1 Study locations

Our three subglacial discharge locations are located in southwest (Fig. 2a, b) and southern (Fig. 2a, c) Greenland (pictures given in Supplementary Material). Our sub-Isunnguata watershed (IS; 67°09'27.1" N, 50°03'25.0" W) and Russell Glacier watershed (RU; 67°05'22.1" N, 50°14'18.7" W) drain to the Akuliarusiarsuup Kuua, which is a tributary to the Qinnguata Kuussua. The short stretch of river downstream of the confluence of the Akuliarusiarsuup Kuua and the main stem of thecQinnguata Kuussua near the town of Kangerlussuaq is also known as the Watson River (Fig. 2b). The majority of drainage from the Isunnguata Glacier drains to the northern Isortoq River (Fig. 2), however a sub-catchment of the Isunnguata drains to a stream that directly feeds the Akuliarusiarsuup Kua, which we refer to as the Northern Tributary, while the Akuliarusiarsuup Kuua refers to river that flows between the outlet of the Russell Glacier and the confluence with the Qinnguata Kuussua (Fig. 2b). Watson River discharge is monitored by PROMICE (Programme for Monitoring of the

Greenland Ice Sheet; van As et al., 2018) and total discharge was 4.3 and 3.6 km^3 of water in 2017 and 2018, respectively (van As et al., 2018). The total catchment size for the Isunnguata is 15,900 km^2, though the size of the sub-catchment draining to the Northern Tributary is smaller with a drainage area of approximately 40 km^2 (Lindbäck et al., 2015; Rennermalm et al., 2013), therefore we hereby refer to this site as the sub-Isunnguata watershed. The total drainage area for the Russell glacier is not precisely known, however the catchment draining both the Russell and Leverett glaciers has been estimated at approximately 900 km^2 (Lindback et al., 2015) while the Leverett drainage area alone is estimated at approximately 600 km^2 (Hawkings et al., 2016), therefore we estimate the Russell drainage area at approximately 300 km^2. Discharge from the third site in southern Greenland, Kiattut Sermiat (KS; 61°12'13.5" N, 45°19'49.1"W) drains to the Kuusuaq River near the town of Narsarsuaq. While Kuusuaq discharge is not monitored, a previous study using dye tracing techniques estimated approximately 0.22 km^3 of discharge in 2013, and its catchment size was estimated at 36 km^2 (Hawkings et al., 2016).

Underlying lithologies differ between sites. Southwest Qinnguata Kuussua sites are located near the boundary between the Archean Craton to the south and the southern Nagssugtoqidian Orogen to the north (Henriksen et al., 2009). The Archean Block is composed of granites and granulite facies orthogneisses that were intruded by mafic dykes during Paleoproterozoic rifting. These rocks were deformed and modified during subsequent continent-to-continent collision in the Paleoproterozoic to create the amphibolite facies gneisses of the southern Nagssugtoqidian Orogen (van Gool et al., 2002). Kiattut Sermiat lies within the Paleoproterozoic Ketilidian fold belt (Henriksen et al., 2009). Lithologies in this region include the Julianehåb Granite and associated basic intrusions and the sedimentary and volcanic rocks of the Mesoproterozic Gardar Province that include a suite of alkaline igneous rocks and basaltic dykes with interbedded sandstones (Kalsbeek and Taylor, 1985; Upton et al., 2003).

Previous studies have characterized chemical weathering reactions in subglacial discharge to the Akuliarusiarsuup Kua and Qinnguata Kuussua (Deuerling et al., 2018; Hasholt et al., 2018; Yde et al., 2014), in the Kuusuaq that drains Kiattut Sermiat (Hawkings et al., 2016), and in comparison between these two sites (Urra et al., 2019). There has been extensive work regarding ice sheet dynamics and hydrology in the southeast glaciers draining to the Akuliarusiarsuup Kua and Qinnguata Kuussua (Van As et al., 2017, 2018; Lindbäck et al., 2015) as well as southern Kuusuaq catchments(Warren and Glasser, 1992; Winsor et al., 2014). Subglacial permafrost has been identified near the sub-Isunnguata site (Ruskeeniemi et al., 2018) and mostly likely formed during Holocene fluctuations in the ice sheet margin. While a similar Holocene ice retreat and re-advance may have occurred in southern Greenland (Larsen et al., 2016), it is unknown whether this retreat led to the organic deposits below the ice sheet.

2.2 Sample collection

We collected water samples from subglacial discharge sites in spring and fall of 2017, and the summer of 2018 to observe seasonal variations in water chemistry. To minimize atmospheric influence, samples were collected as close as possible to the glacier front where subglacial waters emerge, which was less than 10 m for the sub-Isunnguata site and approximately 100 m for the Russell glacier site. Subglacial discharge from Kiattut Sermiat site flowed through a glacial meltwater lake prior

to arriving at the sampling location, which was approximately 1.1 km from the glacial outlet (Fig. 2) and therefore may experience a more interaction with atmospheric gases during transport from the subglacial discharge site to our sampling location. We collected water samples by pumping water through a 0.5-cm flexible PVC tube that was placed in flowing water as far as possible from shore (approximately 1-2 m). A YSI Pro-Plus sensor that was calibrated daily was installed in an overflow cup filled from the bottom to measure specific conductivity (Sp.C), temperature, pH, dissolved oxygen, and oxidation-reduction potential (ORP). These parameters were monitored until stable, between about 10 and 30 minutes, after which samples were collected and preserved in the field according to the solute to be measured after being filtered through a 0.45 µm trace-metal grade Geotech high-capacity disposable canister filter. Samples for cations and anions were collected in HDPE bottles; cation samples were preserved with Optima-grade ultrapure nitric acid (pH<2) while no preservative was added to anion samples. Dissolved inorganic carbon (DIC) samples were filtered through 0.2 µm filters directly to the bottom of 20 ml Qorpac glass vials and allowed to overflow until sealed tightly with no headspace.

Gas samples were collected in duplicate via headspace extractions according to methods outlined in Repo et al. (2007) and Pain et al. (2019). Unfiltered water was pumped into the bottom of 500 mL bottles until they overflowed. Bottles were immediately capped with rubber stoppers fitted with two 3-way inlet valves. 60 mL of water was extracted from one inlet and replaced with 60 mL of atmospheric air (for spring and fall 2017 sampling trips) or ultrapure N_2 gas in a gas bag (summer 2018 sampling trip). Bottles were shaken for 2 minutes to equilibrate headspace gas with water, and headspace gas was extracted and immediately injected into 60 ml glass serum bottles that had been evacuated immediately prior to sample introduction. Samples were stored at room temperature until analysis, which occurred within one week of collection. Measured headspace concentrations were converted to dissolved concentrations using methods outlined in Pain et al. (2019). When atmospheric air was used for headspace extractions, atmosphere samples were collected in tandem and analyzed to correct the calculated dissolved CO_2 and CH_4 concentrations and isotopic compositions for atmospheric CO_2 and CH_4. This correction altered CH_4 concentrations by up to 22% for one sample from the Russell glacier, though less than 5% for all other samples, and resulted in a correction of $\delta^{13}C\text{-}CH_4$ of up to 1.3‰. For CO_2, the correction altered concentrations by up to 15% for one sample collected at Kiattut Sermiat, though less than 10% for all other samples, and resulted in a correction of $\delta\,^{13}C\text{-}CO_2$ of up to 0.4‰.

In samples collected in fall 2017 and summer 2018, alkalinity was measured in the field laboratory within 3 days of collection by titration with 0.01 N HCl using the Gran method. Because alkalinity measurements were not available for the spring 2017 sampling trip, we estimate alkalinity with PHREEQc modeling and the phreeqc.dat database (Parkhurst, 1997) using major cations and anions, pH, temperature, and DIC concentrations as model inputs.

2.3 Laboratory analysis

Gas samples were analyzed for CO_2 and CH_4 concentrations, and $\delta^{13}C\text{-}CO_2$ and $\delta^{13}C\text{-}CH_4$ on a Picarro G2201-i cavity ring-down spectrometer in the field within a few days of collection. Carbon isotopic compositions are reported in reference to Vienna Pee Dee Belemnite (VPDB). Check standards of known CO_2 and CH_4 concentrations and isotopic compositions were

measured during each sample run and were accurate within 10%. Anion and cation concentrations were measured on an automated Dionex ICS-2100 and ICS-1600 Ion Chromatograph, respectively. Error on replicate analyses was less than 5%. DIC concentrations were measured on a UIC (Coulometrics) 5011 CO_2 coulometer coupled with an AutoMate Preparation Device. Samples were acidified and the evolved CO_2 was carried through a silver nitrate scrubber to the coulometer where total C was measured. Accuracy was calculated to be ±0.1 mg/L based on measurement of check standards.

2.4 Methane modeling

To assess CH_4 sources and sinks, we calculate ε_c, or the carbon isotopic fractionation factor between CO_2 and CH_4 as defined in Whiticar (1999):

$$\varepsilon_c = \delta^{13}C_{CO2} - \delta^{13}C_{CH4} \quad (5)$$

Values of ε_c reflect methanogenesis pathways (acetoclastic or CO_2 reduction) as well as the extent of oxidation. Values of ε_c between approximately 40 and 55‰ are produced for CH_4 generated via acetoclastic methanogenesis, while CO_2 reduction produces values between approximately 55 and 90‰. Lower values (ε_c between 5 and 30) result when CH_4 oxidation predominates. Modern atmospheric input without additional alteration of CO_2 or CH_4 isotopic systematics results in a ε_c value of approximately 40 (Whiticar, 1999).

We calculated CH_4 oxidation using the isotopic method outlined in Mahieu et al. (2008) and Preuss et al. (2013). The fraction of oxidized methane (f_{ox}) in an open system is given by:

$$f_{ox} = \frac{\delta_E - \delta_P}{1000 \cdot (\alpha_{ox} - \alpha_{trans})} \quad (6)$$

where δ_E is the measured δ^{13}C-CH_4 value for each water sample, δ_P is δ^{13}C-CH_4 of produced methane, α_{ox} is the oxidation fractionation factor, and α_{trans} is a fractionation factor resulting from diffusive transportation of CH_4. While the exact value of δ_P is unknown, diagenetic alteration of δ^{13}C-CH_4 values through oxidation or transport only enrich δ^{13}C-CH_4 signatures, therefore the value of δ_P is taken as the most depleted δ^{13}C-CH_4 signature assuming it is the least impacted by diagenetic alteration. Literature-reported values for α_{ox} range between 1.003 and 1.049. We calculate the fraction of oxidized methane with the largest fraction factor (α_{ox} = 1.049; Mahieu et al., 2008), which yields the minimum amount of CH_4 oxidation required to explain the observed variations in $\delta^{13}CH_4$, and thus is a conservative estimate for CH_4 oxidation, and actual oxidation ratios may be higher. Literature-reported values for α_{trans} vary from 1 for advection-dominated systems to 1.0178 for diffusion-dominated porous media (Visscher et al., 2004; Mahieu et al., 2008; Preuss et al., 2013). We assume that transport is advection dominated and thus assume α_{trans} = 1, however diffusive transport of CH_4 may result in fractionation of CH_4 in the subglacial environment and lead to relatively lower estimates of f_{ox}. Because hydrologic connectivity between subglacial methanogenic meltwater pockets and drainage features is not well described, the relative importance of advective compared to diffusive CH_4

transport within the subglacial drainage system is not well understood; however, it is presumed to be an advection-dominated system in which expanding drainage networks access and drain methanogenic meltwater pockets throughout the melt season.

2.5 Mineral weathering and carbonate modeling

We used major cation and anion concentrations and alkalinity to partition solutes into the four mineral weathering reactions in Table 1 after correcting solute concentrations for marine aerosol deposition using measured chloride concentrations and standard seawater element ratios. The mass balance model followed the methods of Deuerling et al. (2019). After apportioning solutes to mineral weathering reactions, we used the stoichiometries of reactions to calculate the impact of each reaction on dissolved CO_2 concentrations (Table 1). The mineral weathering model apportions solutes to reactions in Table 1 based on the ratios of Ca/Na and Mg/Na in silicate minerals in stream bedload samples, which were taken to be 0.54 and 0.38, respectively, for sub-Isunnguata and Russell Glacier samples (Deuerling et al., 2019; Hindshaw et al., 2014; Wimpenny et al., 2010, 2011) and 0.39 and 0.27, respectively, for Kiattut Sermiat samples (Da Prat and Martin, 2019). Because mineral weathering reactions may both add and remove CO_2, we discuss both the net impact of mineral weathering on CO_2 concentrations (Net $CO_{2\text{-MW}}$), which may have a positive or negative value:

$$[\text{Net } CO_{2\text{-MW}}] = [CO_{2\text{-CarbCA}}] + [CO_{2\text{-CarbSA}}] + [CO_{2\text{-SilCA}}] \quad (7)$$

as well as the total impact of mineral weathering on CO_2 concentrations (Total $CO_{2\text{-MW}}$),

$$[\text{Total } CO_{2\text{-MW}}] = |[CO_{2\text{-CarbCA}}]| + |[CO_{2\text{-CarbSA}}]| + |[CO_{2\text{-SilCA}}]| \quad (8)$$

where changes in the concentrations of CO_2 are defined by their absolute values. To discuss the relative importance of individual reactions, we define proportional contributions of each reaction as follows:

$$\%CO_{2\text{-CarbCA}} = \frac{|[CO_{2-CarbCA}]|}{[Total\ CO_{2-MW}]} * 100 \quad (9a)$$

$$\%CO_{2\text{-CarbSA}} = \frac{|[CO_{2-CarbSA}]|}{[Total\ CO_{2-MW}]} * 100 \quad (9b)$$

$$\%CO_{2\text{-SilCA}} = \frac{|[CO_{2-SilCA}]|}{[Total\ CO_{2-MW}]} * 100 \quad (9c)$$

We combine measured CO_2 concentrations with Net $CO_{2\text{-MW}}$ in order to determine the magnitude of CO_2 production or consumption in the subglacial environment due to processes besides mineral weathering. This analysis assumes that the concentration of CO_2 measured at the subglacial outlet is equivalent to the net change in CO_2 due to mineral weathering plus the sum of all other subglacial CO_2 sources and sinks. We refer to the sum of all other subglacial CO_2 sources and sinks as $CO_{2\text{-total}}$, which represents the amount of CO_2 that must have been supplied to the subglacial environment to balance the mineral weathering CO_2 sink:

$$CO_{2\text{-measured}} = \text{Net } CO_{2\text{-MW}} + CO_{2\text{-total}} \quad (10)$$

The sources of CO_2 to $CO_{2\text{-total}}$ may be evaluated through the use of Keeling plots, which are constructed as the inverse of CO_2 concentrations ($[CO_2]^{-1}$) versus stable isotopic composition ($\delta^{13}C\text{-}CO_2$). If variations in the concentration and isotopic composition of CO_2 arise from the mixing of two CO_2 reservoirs with constant isotopic compositions and concentrations (Keeling, 1958), a linear relationship is expected between $[CO_2]^{-1}$ and $\delta^{13}C\text{-}CO_2$. The y-intercept of a regression between these variables represents the isotopic composition of the high-CO_2 end member. Because measured CO_2 concentrations include both subglacial CO_2 sources and sinks, which may include considerable consumption through mineral weathering reactions, the magnitude of the total subglacial CO_2 source is taken as $CO_{2\text{-total}}$. We therefore construct Keeling plots between $[CO_{2\text{-total}}]^{-1}$ and measured $\delta^{13}C\text{-}CO_2$ values because while mineral weathering impacts the concentration of CO_2, its isotopic composition is not appreciably altered (Myrttinen et al., 2012) compared to the range of isotopic compositions of potential CO_2 end members, namely OM remineralization, atmospheric CO_2, and lithogenic CO_2 sources due to mechanical grinding (Fig. 1).

2.6 Discharge relationships

We evaluate the relationship between subglacial CH_4 and CO_2 dynamics and glacial meltwater river discharge records collected downstream of the sub-Isunnguata and upstream of the Russell sampling sites. Proglacial river discharge was collected in the Akuliarusiarsuup Kuua (AK) River at the AK4 site, 2 km downstream of the sub-Isunnguata sampling site (Fig. 2b). The river discharge dataset is an updated and extended version of Rennermalm et al. (2012) using reference and regression models to correct Solinst level logger drift in water stage (Solinst, 2017), and a total of 57 discharge measurements to convert continuous water stage to discharge. The standard uncertainty (i.e. the 68th percent confidence interval or one standard deviation) was determined to be 17% using methods and recommendations provided by (Herschy, 1999; ISO Guide 98-3, 2008; WMO, 2010).

Because diurnal fluctuations in river discharge can be large, and differing water travel times from subglacial outlet sites to the discharge monitoring site induces a lag between maximum daily discharge at subglacial discharge sites and the AK4 site outlet, we compare subglacial CH_4 and CO_2 concentrations to average daily discharge, calculated as the average of hourly discharge estimates over the days on which subglacial discharge water samples were collected. We use the AK4 site discharge records for evaluating concentration-discharge relationships for both sub-Isunnguata and Russell sites. Although this site is upstream of the Russell Glacier, its close proximity to the Russell Glacier suggests it is more likely to reflect local melting patterns similar to those that would be controlling discharge dynamics at the Russell than discharge records collected at the Watson River outlet. While Watson River discharge records are also available through PROMICE (van As et al., 2018), which includes some contributions from the Russell glacier, the Watson River includes discharge from the Akuliarusiarsuup Kuua (draining sub-Isunnguata, Russell, and Leverett catchments) as well as the much larger Qinnguata Kuussua catchment and therefore Watson River discharge records are not likely to be representative of the temporal changes in the magnitude and variability of discharge from the much smaller Russell glacier catchment.

3 Results

3.1 Temporal variability in water chemistry and gas concentrations

Chemical parameters differ spatially between subglacial discharge sites as well as temporally through the 2017-2018 melt seasons. Comparing the means and standard deviations of water samples collected throughout 2017 and 2018, specific conductivity (Sp.C; Fig. 3a) is typically highest at Kiattut Sermiat (26±8 µS/cm), followed by Russell (22±5 µS/cm) and sub-Isunnguata sites (13±9 µS/cm; Fig. 3a). All sites show variability over time, with lowest values occurring in the summer for sub-Isunnguata and Russell, while Sp.C drops continuously with days of the year for Kiattut Sermiat. Sites differ in pH, and average values at Kiattut Sermiat (8.2±0.4) are higher than both Russell (7.2±0.2) and sub-Isunnguata (6.6±0.6; Fig. 3b), and while values vary over time, no consistent trend is identified between sites. The saturation of dissolved oxygen (D.O.) with respect to atmospheric concentrations is similar between sites, though sub-Isunnguata (98±8%) values fall below Russell (115±16%) and Kiattut Sermiat (117±11%) during all sampling times and exhibit undersaturation in the mid-summer samples, while Russell and Kiattut Sermiat are consistently supersaturated (Fig. 3c). Alkalinity is similar at Russell (93±31 µeq/L) and Kiattut Sermiat (93±26 µeq/L) which are higher than at sub-Isunnguata (39±25 µeq/L), but all reach minimum values in summer (Fig. 3d). CH_4 concentrations differ by orders of magnitude between sites (Fig. 3e) and are consistently supersaturated with respect to atmospheric concentrations at sub-Isunnguata (648±411 ppm or 1575±997 nM) and Russell (58±33 ppm or 110±78 nM) sites, while close to atmospheric equilibrium at Kiattut Sermiat (4±2 ppm or 9±5 nM). Mean $\delta^{13}C\text{-}CH_4$ values (Fig. 3f) are similar between sub-Isunnguata (-54.7±7.5‰), Russell (-52±7.3‰), and Kiattut Sermiat (-57.6±14.2‰). Measured CO_2 concentrations (Fig. 3g) are consistently supersaturated with respect to atmospheric concentrations for sub-Isunnguata (685±230 ppm or 58±18 µM), near atmospheric equilibrium for Russell (442±31 ppm or 29±4 µM) and undersaturated for Kiattut Sermiat (263±33 ppm or 19±2 µM). Mean $\delta^{13}C\text{-}CO_2$ values (Fig. 3h) are lower in spring and fall for sub-Isunnguata (-16.6±4.0‰) compared to Russell (-13.7±2.3‰) and Kiattut Sermiat (-16.1±1.6‰) sites, though similar seasonal variation occurs for all sites with relatively more depleted values in the spring and fall compared to summer.

3.2 Methane oxidation and relationship with discharge

Values of ε_c are similar over time for sub-Isunnguata (38±10‰) and Russell (38±9‰) and are relatively higher in the summer sampling period, while Kiattut Sermiat ε_c values are higher on average (42±13‰ with lowest values in the summer (Fig. 4a). Estimates of f_{ox} are similar between sub-Isunnguata (17±15%), Russell (23±15%), and Kiattut Sermiat sites (25±22%; Fig. 4b). However, f_{ox} values are higher in the spring and fall sampling times compared to summer for sub-Isunnguata and Russell and approach 50% in the spring, while Kiattut Sermiat values decrease throughout the melt season.

CH_4 concentrations, $\delta^{13}C\text{-}CH_4$ values, and f_{ox} are negatively correlated to average daily discharge for both sub-Isunnguata and Russell sites (Fig. 5a, b and d), while ε_c is positively correlated with discharge for both sub-Isunnguata and Russell (Fig. 5c).

3.3 Mineral weathering impacts on CO_2 and relationship with discharge

Mineral weathering leads to net sequestration of CO_2 at all three sites (Fig. 6a). The magnitude of Net ΔCO_2 differs between sites with the lowest average values at sub-Isunnguata (-39±37 µM) followed by Russell (-65±32 µM) and Kiattut Sermiat (-98±17 µM) sites. Individual mineral weathering reactions produce differing contributions between sites and over time, with notable differences between southwest sites (sub-Isunnguata and Russell) and the southern Kiattut Sermiat site (Fig. 6b). For instance, the proportional contribution of $Carb_{SA}$ is similar between sub-Isunnguata (17±11%) and Russell (15±6%), but lower at Kiattut Sermiat (8±1%; Fig. 6b). Kiattut Sermiat has a relatively greater contribution from $Carb_{CA}$ (62±2%) compared to sub-Isunnguata (41±10%) and Russell (38±6%), while Sil_{CA} is lower at Kiattut Sermiat (28±1%) compared to sub-Isunnguata (41±17%), and Russell (47±11%). Kiattut Sermiat additionally exhibits low seasonal variability in the proportional contributions of individual mineral weathering reactions compared to sub-Isunnguata and Russell sites.

$CO_{2\text{-total}}$ represents CO_2 concentrations in the subglacial environment prior to addition and/or consumption of CO_2 through mineral weathering (Eq. 10; Fig. 7). Because the Net $CO_{2\text{-MW}}$ is always negative (more consumption than production), the value of $CO_{2\text{-total}}$ is always greater than measured concentrations ($CO_{2\text{-measured}}$). Regardless of differences in $CO_{2\text{-measured}}$ between sites, the average $CO_{2\text{-total}}$ values are similar between sites and average 91±47 µM for sub-Isunnguata, 94±33 µM for Russell, and 117±16 µM for Kiattut Sermiat.

For both sub-Isunnguata and Russell sites, average daily discharge is negatively correlated with CO_2 concentrations (Fig. 8a), while positively correlated with $\delta^{13}C\text{-}CO_2$ (Fig. 8b).

Keeling plots between $[CO_{2\text{-total}}]^{-1}$ and $\delta^{13}C\text{-}CO_2$ for each site indicate no linear relationship for Russell or Kiattut Sermiat samples, however a strong linear correlation is observed for sub-Isunnguata (r^2=0.99; $p<0.001$) samples with the removal of one outlier, which also had the lowest $CO_{2\text{-total}}$ value (Fig 9).

4 Discussion

We observe orders of magnitude variability in dissolved CH_4 and CO_2 concentrations in subglacial discharge of the Greenland Ice Sheet, indicating significant differences in the magnitudes of the sources and sinks of these gases across time and space. Supersaturation of both CO_2 and CH_4 with respect to atmospheric concentrations indicates that sub-Isunnguata discharge is a source of both gases to the atmosphere, neighbouring Russell Glacier discharges water that is a source of CH_4 but near equilibrium with respect to CO_2, while Kiattut Sermiat in southern Greenland is a sink of atmospheric CO_2 but near equilibrium with respect to CH_4 (Fig. 3e, g). Because CH_4 dynamics may be largely microbially driven while CO_2 dynamics include microbial as well as abiotic mineral weathering processes, we first discuss CH_4 dynamics including a comparison of concentrations, isotopic compositions, and extent of oxidation between sites and over the melt season. We then discuss CO_2 concentrations, impacts of mineral weathering reactions (Table 1), and an assessment of subglacial CO_2 sources, including

OM remineralization. These assessments will contribute to our understanding of the variability and controls of CH_4 and CO_2 concentrations in subglacial discharge from the Greenland Ice Sheet and may improve predictions of the impact of future ice melt on Arctic carbon budgets.

4.1 Sources and sinks of CH_4

Differences in CH_4 concentrations and relationships with discharge between sites imply heterogeneity in both the extent and controls of subglacial methanogenesis under the Greenland Ice Sheet. CH_4 supersaturation occurs at the two subglacial discharge sites that flow to the Akuliarusiarsuup Kuua (sub-Isunnguata and Russell), and concentrations are similar to the ranges reported in discharge of the Leverett Glacier (up to 600 nM; Lamarche-Gagnon et al., 2019), located near the Russell Glacier in this study (Fig. 2b). However, CH_4 concentrations are near atmospheric equilibrium for the Kiattut Sermiat site (Fig 3e). Because methanogenesis is an anaerobic OM remineralization pathway, it is more likely to occur in subglacial environments isolated from atmospheric O_2 sources. Widespread observations of methanogenesis in glacial meltwater of southwest Greenland from this and other studies (Christiansen and Jørgensen, 2018; Dieser et al., 2014; Lamarche-Gagnon et al., 2019), and limited observations of CH_4 in subglacial discharge in southern Greenland, suggests heterogeneity in subglacial conditions that support methanogenesis. Methanogenesis fuelled by organic material overridden during ice sheet growth has been suggested as a potential climate feedback over glacial interglacial timescales (Wadham et al., 2008), and may contribute to observed variations in CH_4 concentrations.

Subglacial methane concentrations may additionally be controlled by hydrologic factors as the subglacial hydrological network develops throughout the melt season and channelization of meltwater conduits increases subglacial drainage efficiency (Andrews et al., 2015; Cowton et al., 2013). Drainage efficiency impacts both subglacial water residence time as well the transport of aerobic supraglacial meltwater to the ice bed. Both residence time and oxygen delivery may impact subglacial redox status and methanogenesis potential and favour methanogenesis when oxygen supply rates are low compared to OM remineralization rates. This condition is most likely to be met in distributed subglacial systems that are hydrologically isolated with limited inputs from aerobic supraglacial meltwater. Such a hydrologic control on methanogenesis is supported by the negative correlation between CH_4 concentrations and average daily discharge at both sites (Fig. 5a). This correlation would suggest that either CH_4 production occurs predominantly during periods of low discharge and greater residence time, or higher discharge results in the dilution of a CH_4-laden subglacial water source. While both mechanisms would lead to a similar relationship between discharge and CH_4 concentrations, they carry different implications for subglacial methanogenesis. If limited by residence time, a hydrologic link between glacial hydrology and subglacial biogeochemistry would be established due to the importance of supraglacial discharge for the delivery of terminal electron acceptors to the ice bed for the active production of CH_4. If predominantly controlled by dilution, however, active methanogenesis would not be required; only the existence of meltwater pockets containing CH_4 that may or may not have been recently produced. Further discussion of these mechanisms is outside the scope of this study. While we have limited data to make further inferences about

hydrologic controls of methanogenesis, the presence of several outliers at the sub-Isunnguata site in particular (Fig. 5a) highlights the possibility for additional controls including stochastic drainage events or heterogeneity in subglacial CH_4 concentrations that result in variability in the relationship between concentration and discharge, as was observed in Lamarche-Gagnon et al. (2019).

While our results suggest heterogeneity in the extent of methanogenesis between outlet glaciers, they suggest homogeneity of the microbial methanogenesis pathway as well as CH_4 oxidation dynamics between sites. Methanogenesis pathways may be evaluated by $\delta^{13}C\text{-}CH_4$ as well as ε_c values because they impart distinct $\delta^{13}C$ signatures to CH_4 and CO_2 (Whiticar and Schoell, 1986). Dieser et al. (2014) measured a microbial $\delta^{13}C\text{-}CH_4$ production signal at the Russell Glacier with values between -63‰ and -64‰, which was interpreted to reflect a possible combination of CH_4 produced through both acetoclastic and CO_2 reduction pathways. The most depleted $\delta^{13}C\text{-}CH_4$ value measured at the sub-Isunnguata in this study was -62.7‰, close to values measured by Dieser et al. (2014) (Fig. 3f), and similar to values reported by Lamarche-Gagnon et al. (2019) for the Leverett Glacier. The similar isotopic ratio between our samples and that measured in active methanogenic communities suggests similar methanogenesis pathways occur across this region, and additionally suggests that the CH_4 we measure in the sub-Isunnguata comes from active methanogenesis rather than old preserved subglacial CH_4 reservoirs, at least in the peak melt season when we observe these depleted $\delta^{13}C\text{-}CH_4$ values.

While the exact contributions from each methanogenesis pathway cannot be inferred from isotopic information alone, the range of ε_c values at outlet glaciers are consistent with predominantly acetoclastic methanogenesis during the peak melt season (Fig. 4a). However, ε_c values fall below the expected range from acetoclastic methanogenesis during the early and late melt seasons and may result from variations in the extent of subglacial CH_4 oxidation. Seasonal variation in CH_4 oxidation is supported by consistency between ε_c and f_{ox} values, which both indicate the greatest impact of oxidation (approaching 50%) in the early melt season compared to peak melt season (Fig. 4a, b), with additional evidence of elevated CH_4 oxidation in the late melt season at both sub-Isunnguata and Russell sites. Because our water sampling locations were slightly downstream of glacial discharge outlets, there is also the possibility that outgassing in between the outlet and our sampling location reduced dissolved CH_4 concentrations and led to more enriched isotopic compositions of remaining dissolved CH_4. While it is likely that some outgassing did occur, it is unlikely that the extent of outgassing between the glacial outlet and our sampling location would vary significantly between sampling times, and thus outgassing would not explain temporal differences in concentration, $\delta^{13}C\text{-}CH_4$, f_{ox}, or ε_c. While our measured gas concentrations and isotopic compositions likely reflect some modification of CH_4 and CO_2 isotopic compositions due to outgassing, the differences over time are more likely due to changes in subglacial CH_4 dynamics than outgassing.

The extent of CH_4 oxidation may be controlled by multiple factors including oxygen availability, subglacial residences time, and subglacial hydrology, similar to methanogenesis. A hydrologic control of CH_4 oxidation is supported by relationships between f_{ox} and ε_c with average daily discharge (collected at site AK4; Fig. 2b) at both sub-Isunnguata and Russell

sites: f_{ox} is negatively related with discharge for both sites (Fig. 5b) while ε_c is positively correlated with discharge (Fig. 5c). These correlations suggest that CH_4 oxidation is greatest during periods of low flow, which may be associated with longer residence times to allow subglacial CH_4 oxidation, however this relationship could also result from differences in CH_4 sources throughout the melt season as the subglacial drainage network expands. Assuming the former, the delivery of oxygen to the subsurface by supraglacial melting does not appear to be a limiting factor in subglacial CH_4 oxidation, which should increase f_{ox} as more oxygenated supraglacial water is delivered to the subglacial system. Instead, the observed greater CH_4 oxidation during periods of low discharge may reflect mixing of methane-rich subglacial meltwater pockets and aerobic subglacial meltwater leading to CH_4 oxidation. Longer transit times during periods of low flow may allow more subglacial methane oxidation to occur than during peak discharge, when the development of channelized flow paths reduces meltwater residence time in the subglacial environment.

Our results indicate a high degree of heterogeneity in subglacial methanogenesis under the Greenland Ice Sheet, as well as a significant impact of CH_4 oxidation, which serves to reduce atmospheric CH_4 fluxes. Given the observed spatial and temporal heterogeneity of CH_4 concentrations and processes, further investigation of the spatial variability in outlet glacier CH_4 concentrations is needed to determine the impact of Greenland Ice Sheet loss on Arctic and global CH_4 budgets, while a better understanding of the controls of these differences will improve models of how CH_4 fluxes from subglacial discharge will change with continued warming.

4.2 Sources and sinks of CO_2

Dissolved CO_2 concentrations in subglacial discharge are consistently supersaturated with respect to atmospheric concentrations at the sub-Isunnguata site, near atmospheric equilibrium at the Russell Glacier, and undersaturated at Kiattut Sermiat Glacier, indicating that glacial meltwater from the Greenland Ice Sheet can serve as either a source or sink of CO_2 to the atmosphere. Similar to CH_4, differences in dissolved CO_2 dynamics (Fig. 3g) imply variability in carbon processes under the Greenland Ice Sheet. We first discuss potential subglacial CO_2 sources, including OM remineralization, followed by a discussion of CO_2 consumption due to mineral weathering.

4.2.1 Subglacial CO_2 sources

There are many potential sources of CO_2 in the subglacial environment including dissolution of atmospheric gases in air-filled conduits or fractures in ice, CO_2 contained in ice bubbles (Fig. 1; Anklin et al., 1995; Graly et al., 2017) mechanical grinding and volatilization of fluid inclusions in bedrock (Macdonald et al., 2018), and OM remineralization. While previous studies have indicated that additional atmospheric CO_2 input through fractures and air-filled conduits may supply sufficient CO_2 to drive mineral weathering observed in many subglacial environments, including several sites in Greenland (Graly et al., 2017), CO_2 is also a product of OM remineralization, which is believed to account for CH_4 concentrations elevated above atmospheric equilibrium at the two southwest sites in this study. Both CO_2 and CH_4 exhibit negative correlations with average

daily discharge for both sub-Isunnguata and Russell sites, and could suggest a common OM remineralization source (Fig. 8a). While the magnitude of this source and its relative importance compared to other subglacial CO_2 sources is currently unknown, differing sources of carbonic acid for mineral weathering reactions carry different implications for subglacial CO_2 budgets. For instance, carbonic acid weathering driven by invasion of atmospheric CO_2 would represent a sink of atmospheric CO_2, but carbonic acid weathering driven by OM remineralization would instead serve to consume CO_2 from *in situ* sources and limit the potential for subglacial meltwater to be an atmospheric CO_2 source once discharged from the glacier. Determining the sources of carbonic acid to subglacial weathering reactions is therefore critical to understand the controls of mineral weathering in subglacial environments and its role in atmospheric CO_2 sequestration.

Comparisons between measured $\delta^{13}C\text{-}CO_2$ in subglacial discharge samples and likely $\delta^{13}C\text{-}CO_2$ values of CO_2 sources indicate that CO_2 sources differ between sites, with OM remineralization as the most important CO_2 source at the sub-Isunnguata but likely not the predominant or sole source at Russell or Kiattut Sermiat glaciers. Keeling plots of $[CO_{2\text{-total}}]^{-1}$ versus $\delta^{13}C\text{-}CO_2$ indicate that $CO_{2\text{-total}}$ may be represented by a two-end member mixing model for sub-Isunnguata discharge, but not for discharge from the Russell and Kiattut Sermiat glaciers (Fig. 9). Mixing model end members include a ^{13}C-enriched, lower concentration CO_2 source and a ^{13}C-depleted, higher concentration CO_2 source. The y-intercept of the regression between $[CO_{2\text{-total}}]^{-1}$ versus $\delta^{13}C\text{-}CO_2$, which represents the isotopic signature of the high-CO_2 endmember, is -27.4‰. This value is close to what would be expected from OM remineralization as indicated by remineralized OM in Greenlandic heath soils that produced $\delta^{13}C\text{-}CO_2$ of approximately -27 to -25‰ (Ravn et al., 2020), and thawed Alaskan permafrost soils that produced $\delta^{13}C\text{-}CO_2$ of between -20 to -30‰ for (Mauritz et al., 2019), both of which may be similar to subglacial organic matter. The low-CO_2 end member could reflect atmospheric CO_2 input, with an associated $\delta^{13}C\text{-}CO_2$ value of approximately -8‰. While the $\delta^{13}C\text{-}CO_2$ value of the lowest-$CO_{2\text{-total}}$ samples in the sub-Isunnguata Keeling plot (e.g. highest $[CO_{2\text{-total}}]^{-1}$ not including the outlier) is slightly depleted compared to atmospheric values at -12.1‰; however, even the lowest CO_2 concentrations measured at sub-Isunnguata are supersaturated with respect to atmospheric concentrations (Fig. 3g). Thereby suggesting that OM remineralization contributes CO_2 even for low CO_2-concentration samples and isotopically depletes the subglacial CO_2 reservoir.

While $\delta^{13}C\text{-}CO_2$ values of Russell and Kiattut Sermiat samples are within the range of sub-Isunnguata samples, suggesting possible contributions of $CO_{2\text{-atm}}$ and $CO_{2\text{-OM}}$, scatter in the Keeling plots indicate variability in the CO_2 concentration and/or isotopic composition of end members, or significant contributions of at least one other major subglacial CO_2 source. We address both possibilities here. While atmospheric CO_2 concentrations and $\delta^{13}C$ values should be relatively invariable, $CO_{2\text{-OM}}$ may vary both in concentration and isotopic composition, depending on variability in the quantity and composition of organic deposits as well as remineralization rates. For instance, if remineralization largely occurs in hydrologically isolated subglacial meltwater pockets, some variability in the concentration and $\delta^{13}C\text{-}CO_2$ of $CO_{2\text{-OM}}$ is likely. While no data yet exist to characterize the variability in subglacial OM reservoirs, variability in either concentration or isotopic

composition of $CO_{2\text{-}OM}$ could plausibly result in the scatter shown in Figure 9. Additional subglacial CO_2 sources from ice bubbles, or lithogenic CO_2 liberated by mechanical grinding would be expected to enrich rather than deplete the $\delta^{13}C\text{-}CO_2$ values of the samples relative to modern atmospheric $\delta^{13}C\text{-}CO_2$ values. Ice bubbles contain gaseous CO_2 at concentrations and isotopic compositions reflecting atmospheric conditions during ice formation. While heterogeneity may result from gas bubbles recording changes in atmospheric CO_2, variability in $\delta^{13}C\text{-}CO_2$ of gas bubble CO_2 should be only a few per mil, which is small compared to the variation observed in Russell and Kiattut Sermiat samples (e.g., Tipple et al., 2010; Fig. 9). Gas bubble CO_2 should also be ^{13}C-enriched compared to modern atmospheric CO_2 due to fossil fuel contributions, and thus would be unlikely to cause the observed depletion of ^{13}C in the subglacial water.

Recent work has also highlighted the potential for subglacial mechanical grinding to result in CO_2 production through the volatilization of bedrock fluid inclusions (Macdonald et al., 2018). While this process was found to produce sufficient CO_2 to drive approximately 20% of mineral weathering in Svalbard subglacial environments, the expected isotopic composition of lithogenic CO_2 is more ^{13}C-enriched than our measured $\delta^{13}C\text{-}CO_2$ values. For instance, estimates of $\delta^{13}C$ for bulk hydrocarbons in fluid inclusions in the Ilímaussaq alkaline complex of South Greenland have values of -4.5±1.5‰ (Madsen, 2001), which is close to the $\delta^{13}C\text{-}CO_2$ of CO_2 in fluid inclusions in the Bamble granulite sector of South Norway (~ -6‰; Newton et al., 1980). Because mechanical grinding should not fractionate the $\delta^{13}C\text{-}CO_2$ values (Lüders et al., 2012), our low $\delta^{13}C\text{-}CO_2$ values suggest this source is small relative to other sources.

One additional source or sink of CO_2 to some of our samples is atmospheric exchange as water flows from the subglacial outlet site to our sampling sites. However, atmospheric CO_2 exchange after discharge would have the same impact on Keeling plots as atmospheric CO_2 exchange prior to discharge. Incorporation of an atmospheric source between the outlet and sample site would be most likely at Kiattut Sermiat where CO_2 concentrations are undersaturated with respect to atmospheric concentrations, which would promote invasion of atmospheric CO_2. However, the measured $\delta^{13}C\text{-}CO_2$ values are more ^{13}C-depleted than modern atmospheric CO_2 and are not consistent with atmospheric CO_2 as the sole or dominant source of CO_2 to these samples (Fig. 9).

While more information is needed to determine all possible sources of CO_2 to Russell and Kiattut Sermiat samples, $\delta^{13}C\text{-}CO_2$ values of samples from both sites imply mixing between a ^{13}C-depleted CO_2 source, such as OM remineralization, and one or more ^{13}C-enriched CO_2 sources, such as atmospheric or lithogenic CO_2. Similar to CH_4, concentrations and isotopic compositions of gases may be impacted by atmospheric exchange between the glacial outlet and our sampling locations, which would alter dissolved CO_2 concentrations and $\delta^{13}C\text{-}CO_2$ compositions to values more similar to atmospheric values. Therefore, we are unable to distinguish the impacts of atmospheric exchange that occurs prior to discharge from exchange that occurs between discharge and our sampling locations. However, the impact of this exchange should be relatively constant between sampling times and sampling locations, therefore outgassing would not account for temporal or spatial variability in CO_2 concentrations or isotopic compositions between sites.

4.2.2 Subglacial CO_2 sink: mineral weathering reactions

Mineral weathering leads to net CO_2 consumption in all subglacial discharge samples (Fig. 6), and thus the measured CO_2 concentrations at glacial outlets represent only a fraction of the total CO_2 that would have been present in the absence of mineral weathering reactions ($CO_{2\text{-total}}$; Eq. 10). Net consumption occurs because the CO_2 source from $Carb_{SA}$ is ubiquitously low compared to sinks from either $Carb_{CA}$ or Sil_{CA} (Fig. 6b). The range in Net $CO_{2\text{-MW}}$ is similar between subglacial discharge sites (between 10-150 µM; Fig. 6a), but average values increase from Kiattut Sermiat to Russell to sub-Isunnguata, likely reflecting the relative weatherability of alkaline igneous rocks, granulite facies gneisses, and amphibolite facies gneisses. Kiattut Sermiat is characterized by a relatively high proportion of $Carb_{CA}$ compared to sub-Isunnguata and Russell sites, which may arise from the presence of trace carbonates in abundant and readily weatherable basaltic intrusions as has been implicated in other studies (Urra et al., 2019). The relatively greater influence of carbonate dissolution compared to silicate dissolution on Total $CO_{2\text{-MW}}$ at Kiattut Sermiat may also relate to more rapid dissolution kinetics of carbonates, which allow carbonate dissolution to have a large influence on major cation and anion load even when carbonates are only present in trace amounts (Deuerling et al., 2019; Tranter, 2005). At sub-Isunnguata and Russell sites, Sil_{CA} has a greater influence than $Carb_{CA}$ on Total $CO_{2\text{-MW}}$, which could result from either a lower abundance of trace carbonates to participate in weathering reactions, or relatively longer subglacial residence times that would allow a greater accumulation of silicate weathering products.

Despite the impact of $Carb_{CA}$ on Total $CO_{2\text{-MW}}$ at Kiattut Sermiat compared to sub-Isunnguata and Russell sites, $Carb_{SA}$ is notably lower at Kiattut Sermiat than other sites and suggests a limited role for sulfuric acid weathering that may relate to subglacial sulfide oxidation dynamics. Lower abundances of sulfide minerals in the subglacial environment may limit the production of sulfuric acid, and could result from differences in lithology between sites, the depletion of sulfide minerals due to prior weathering (Graly et al., 2014), or weathering occurring in anoxic environments that limit the oxidation of sulfide to sulfuric acid (Deuerling et al., 2019). The kinetics of sulfide oxidation may also significantly differ between sites depending on the relative contributions of abiotic compared to microbially mediated sulfide oxidation, as microbially mediated sulfide oxidation is several orders of magnitude faster than abiotic sulfide oxidation. Rapid microbially mediated sulfide oxidation has been implicated in the development of anaerobic conditions, which could also support subglacial methanogenesis (Sharp et al., 1999). Observations of higher CH_4 concentrations as well as higher contributions of $Carb_{SA}$ at sub-Isunnguata and Russell compared to Kiattut Sermiat may therefore be linked to subglacial microbial activity, which is known to vary based on factors such as the presence of organic and fine-grained rock flour to serve as growth substrates, insulation from fluctuations in temperature, and delivery of nutrients and organic matter from supraglacial sources (Sharp et al., 1999). If microbially driven, our results suggest possible linkages between microbial processes and subglacial mineral weathering regimes, with significant impacts to both CH_4 and CO_2 dynamics due to the role of $Carb_{SA}$ as a CO_2 source (Table 1).

Whether controlled by geochemical, microbial, or mechanical processes, the relationships between CO_2 concentrations and$\delta^{13}C\text{-}CO_2$ and average daily discharge are similar between sub-Isunnguata and Russell sites (Fig. 8). These similarities suggest that underlying controls in carbonate chemistry may be consistent between sites despite the heterogeneity

in measured dissolved CO_2 concentrations. For both sub-Isunnguata and Russell sites, mineral weathering reactions consume CO_2, which implies contributions from *in situ* CO_2 sources (such as atmospheric CO_2 invasion or OM remineralization) to produce measured CO_2 concentrations. The different CO_2 concentrations observed between sites therefore appear to result from the strength of *in situ* CO_2 sources relative to $CO_{2\text{-}MW}$, both of which impart the greatest chemical change at times of low discharge and high subglacial residence time. At Kiattut Sermiat, where measured CO_2 concentrations are lowest, the magnitude of *in situ* sources is insufficient to maintain atmospheric equilibrium, leading subglacial discharge to be a sink of atmospheric CO_2, while $CO_{2\text{-}total}$ maintains close to atmospheric equilibrium concentrations at the Russell Glacier. At sub-Isunnguata, however, OM remineralization produces more CO_2 than is consumed by mineral weathering and causes meltwater to be a source of CO_2 to the atmosphere. This finding implies that subglacial mineral weathering serves to partially or fully consume CO_2 produced from *in situ* sources under the Greenland Ice Sheet but does not always serve to directly consume modern atmospheric CO_2.

5 Conclusions

Subglacial reactions impact the concentrations of CO_2 and CH_4 in subglacial discharge of Greenland Ice Sheet, and differences in the relative magnitudes of microbial and geochemical processes result in a high degree of previously unrecognized heterogeneity between glacial discharge sites of the Greenland Ice Sheet. Our results imply a significant role of OM remineralization in driving this heterogeneity and leading to CO_2 and CH_4 supersaturation at the sub-Isunnguata site, and CH_4 supersaturation at the Russell site. Heterogeneity may result in significant uncertainty in total greenhouse gas flux estimates from subglacial systems of the Greenland Ice Sheet, which will be an increasingly important carbon flux as the Arctic warms in the coming decades. While heterogeneous, the uncertainty in greenhouse gas fluxes from Greenland Ice Sheet meltwater may be reduced by a better understanding of the controls and variability of the weathering reactions and microbial processes driving heterogeneous gas concentrations. Such a process-based understanding could also improve estimates of the impact of greenhouse gas variability associated with the growth and retreat of continental ice sheets on over glacial-interglacial timescales. Subglacial OM remineralization further implies the existence of links between subglacial OM deposits and export of other biogeochemical solutes from the Greenland Ice Sheet, including nutrients as well as redox-sensitive elements. While the export of nutrients from the Greenland Ice Sheet has been the focus of numerous studies (Bhatia et al., 2013; Hawkings et al., 2016; Lawson et al., 2014), little is currently known regarding the role of OM sources in governing these exports. Given the variability in greenhouse gas concentrations observed in this study, constraining the extent of heterogeneity in outlet glaciers of the Greenland Ice Sheet as well as the biogeochemical, hydrologic, and geologic controls of this heterogeneity will be important for upscaling atmospheric fluxes as well as efforts to predict impacts of ice loss on carbon and nutrient budgets due to current and future melting of the Greenland Ice Sheet.

Author Contribution

Jonathan B. Martin and Ellen E. Martin participated in conceptualization, data collection, data interpretation, reviewing and editing the manuscript, and acquired the funding for this project. Asa Rennermalm provided the discharge data that was used in analyses of discharge-concentration relationships. Shaily Rahman participated in data collection and early interpretation and presentation of results. Andrea J. Pain participated in conceptualization, data collection, analysis, and interpretation, and took the lead on writing the manuscript with contributions from Jonathan B. Martin, Ellen E. Martin, and Asa Rennermalm.

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Table 1. Mineral weathering reactions and impacts on dissolved CO_2 concentrations.

Eq.	*Mineral*	*Acid*	*Abbreviation*[1]	*Reaction*	*Impact on CO_2*
(1)	Carbonate	Carbonic	$Carb_{CA}$	$(Ca, Mg)CO_3 + H_2O + \mathbf{CO_2} \rightarrow (Ca^{2+}, Mg^{2+}) + 2HCO_3^-$	CO_2 sink
(2)		Sulfuric	$Carb_{SA}$	$2(Ca, Mg)CO_3 + H_2SO_4 \rightarrow 2(Ca^{2+}, Mg^{2+}) + SO_4^{-2} + H_2O + \mathbf{CO_2}$	CO_2 source
(3a)	Silicate	Carbonic	Sil_{CA}	$(Ca, Mg)Al_2Si_2O_8 + \mathbf{2CO_2} + 3H_2O \rightarrow (Ca^{2+}, Mg^{2+}) + 2HCO_3^- + Al_2Si_2O_5(OH)_4$	CO_2 sink
(3b)				$(Na, K)AlSi_3O_8 + \mathbf{CO_2} + 5.5H_2O \rightarrow (Na, K) + HCO_3^- + 0.5Al_2Si_2O_5(OH)_4 + 2H_4SiO_4$	CO_2 sink
(4a)		Sulfuric	Sil_{SA}	$(Ca, Mg)Al_2Si_2O_8 + H_2SO_4 + H2O \rightarrow (Ca^{2+}, Mg^{2+}) + SO_4^{-2} + Al_2Si_2O_5(OH)_4$	No impact
(4b)				$2(Na, K)AlSi_3O_8 + H_2SO_4 + 9H_2O \rightarrow 2(Na^+, K^+) + SO_4^{2-} + Al_2Si_2O_5(OH)_4$	No impact

[1] Abbreviations are based first on the mineral class (carbonate = carb, silicate = sil) and then on the acid (carbonic acid = CA, sulfuric acid = SA)

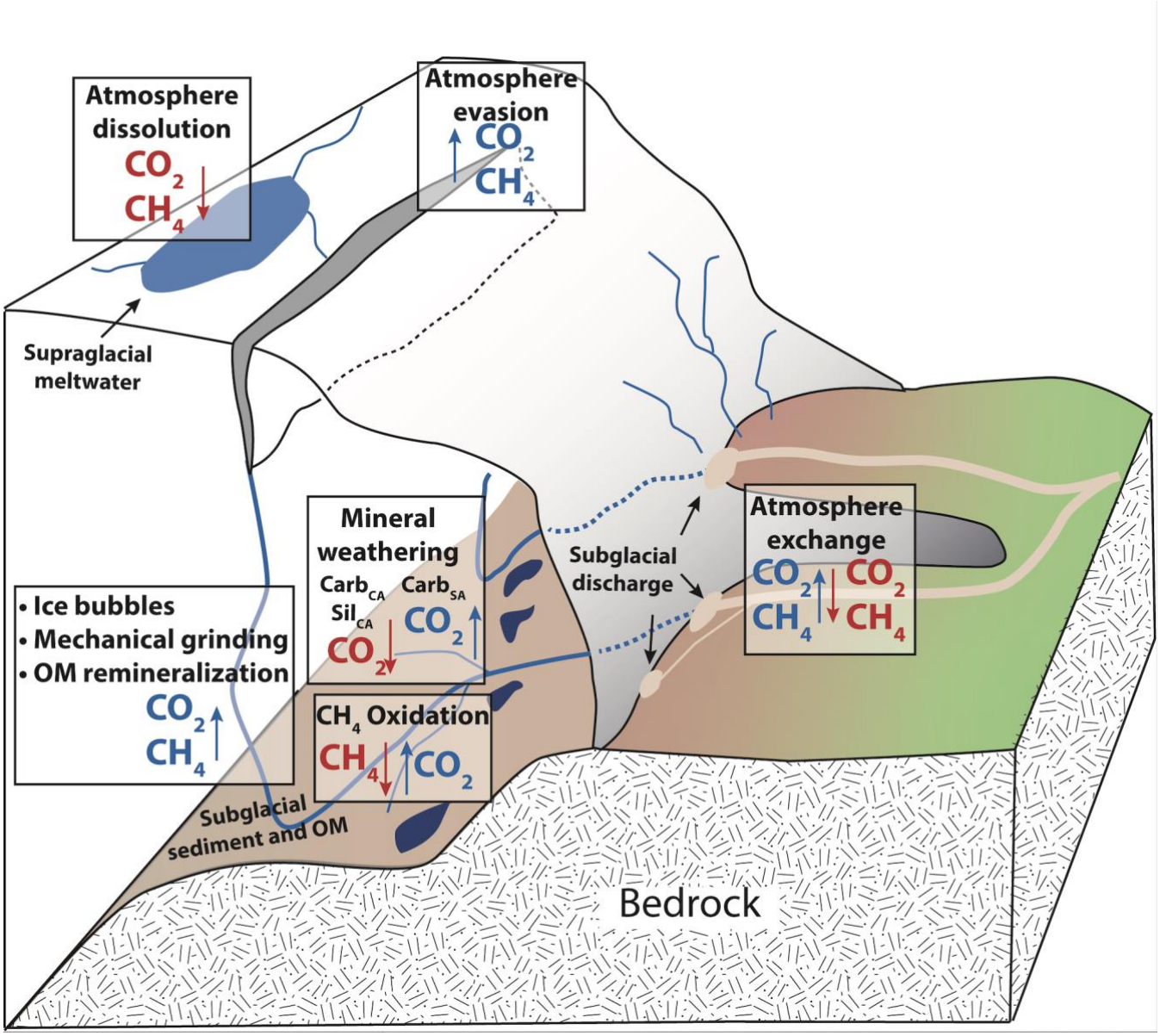

Figure 1: Conceptual diagram of subglacial sources and sinks of CO_2 and CH_4. Arrows indicate the direction of fluxes. Boxes represent processes, and sources of gases to subglacial meltwaters are indicated by blue text while sinks of gases to subglacial meltwater are indicated by red text. Gas bubbles, mechanical grinding, and OM remineralization are grouped because all are CO_2 and CH_4 sources.

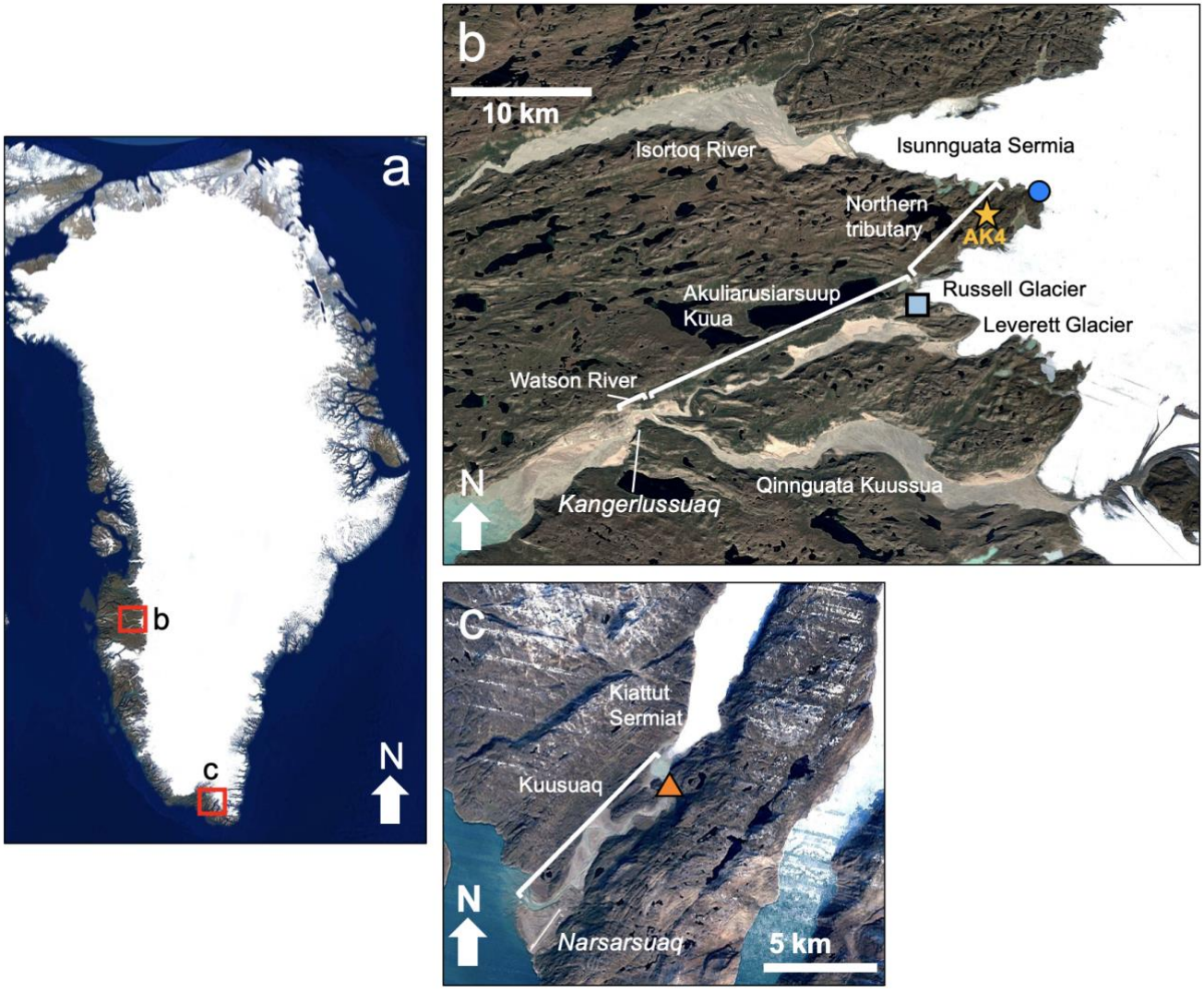

Figure 2. Google Earth satellite images of study locations in a) Greenland including b) locations near the town of Kangerlussuaq, including the sub-Isunnguata water sampling location (dark blue circle) and Russell water sampling location (light blue square). The gold star represents the location of AK4, where proglacial river discharge records were collected. (c) Location of Kiattut Sermiat site (orange triangle) near the town of Narsarsuaq in southern Greenland where water samples were collected.

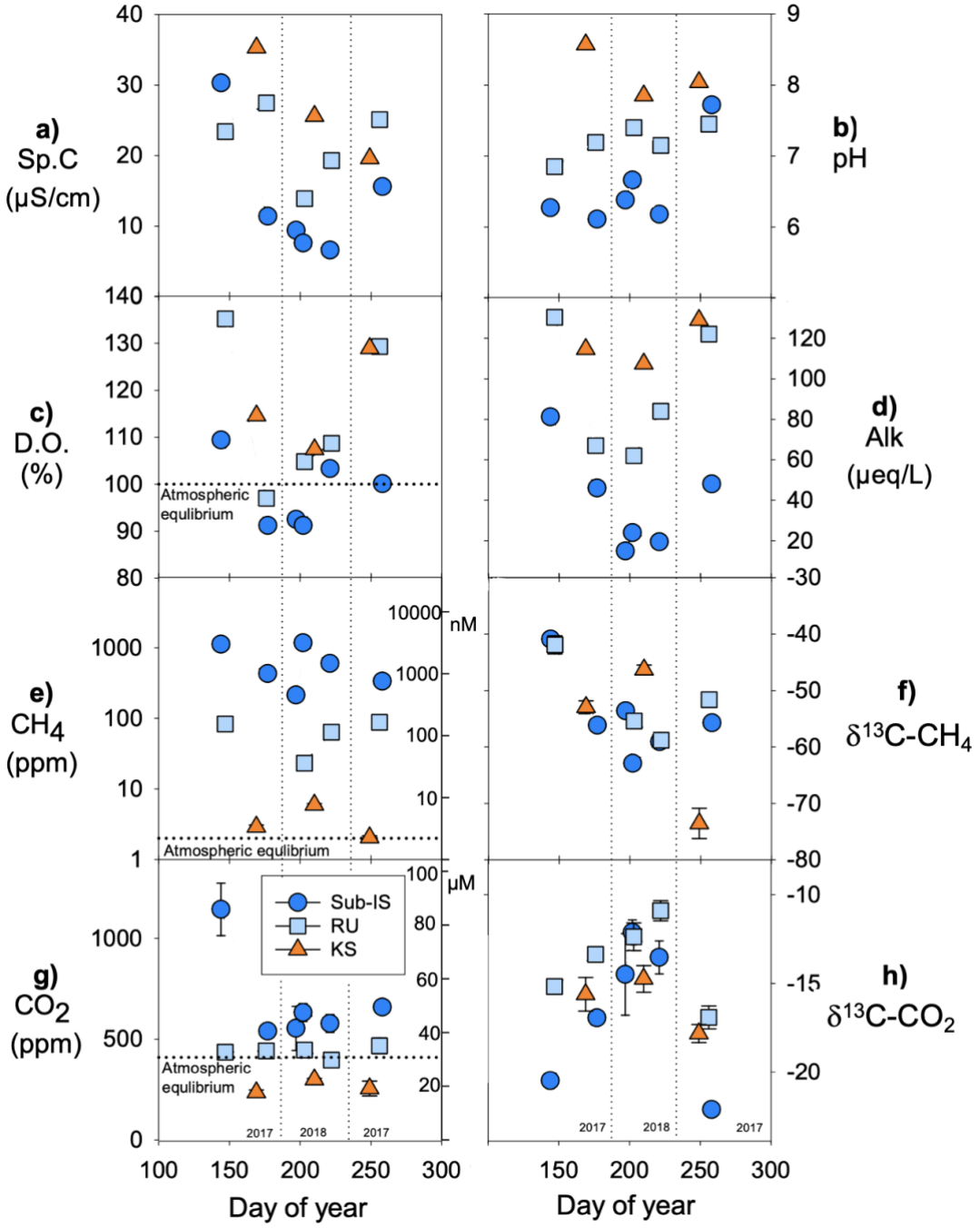

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Figure 3. Chemical parameters at Isunnguata (IS), Russell (RU) and Kiattut Sermiat (KS) subglacial water sampling sites versus day of year for a) specific conductivity (Sp.C), b) pH, c) dissolved oxygen (D.O.) percent saturation, d) alkalinity (Alk), e) measured CH_4 concentrations (left y-axis in ppm and right y-axis in nM), f) δ^{13}-CH_4 values, g) measured CO_2 concentrations (left y-axis in ppm and right y-axis in µM), and h) $\delta^{13}C$-CO_2 values. Atmospheric equilibrium concentrations are indicated by dashed lines and
735 **taken as 1.9 ppm for CH_4 and 410 ppm for CO_2. Error bars on CH_4 and CO_2 concentrations and stable isotopic compositions represents the standard deviation of replicates and are smaller than symbols for some data points.**

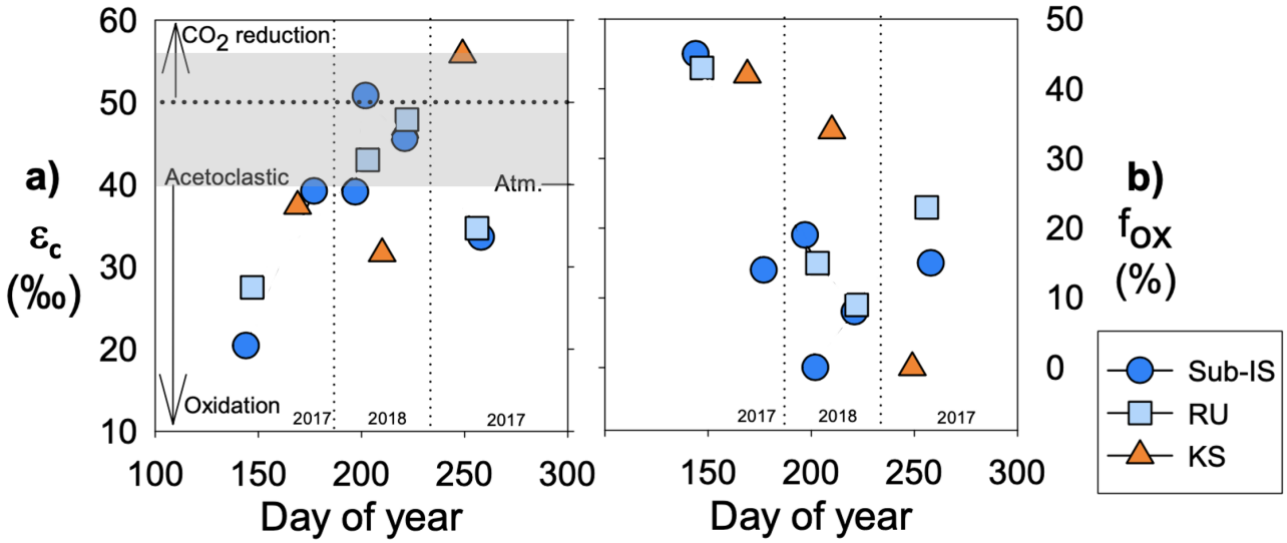

Figure 4. CH_4 dynamics over the course of the 2017 and 2018 melt seasons including a) the carbon fractionation factor (ε_c) between dissolved CO_2 and CH_4 and b) the fraction of CH_4 oxidized (fo_x) for sub-Isunnguata (IS), Russell (RU) and Kiattut Sermiat (KS) samples. Fields of ε_c representing methanogenesis and oxidation values are based on Whiticar (1999). Values of ε_c between approximately 40 and 55 (gray shaded region in panel a) are produced for methanogenesis via acetate fermentation, while CO_2 reduction produces values between approximately 50 and 90. Lower values result from a predominant isotopic signature of CH_4 oxidation. Atmospheric input without additional alteration of CO_2 or CH_4 isotopic systematics results in a ε_c value of approximately 40.

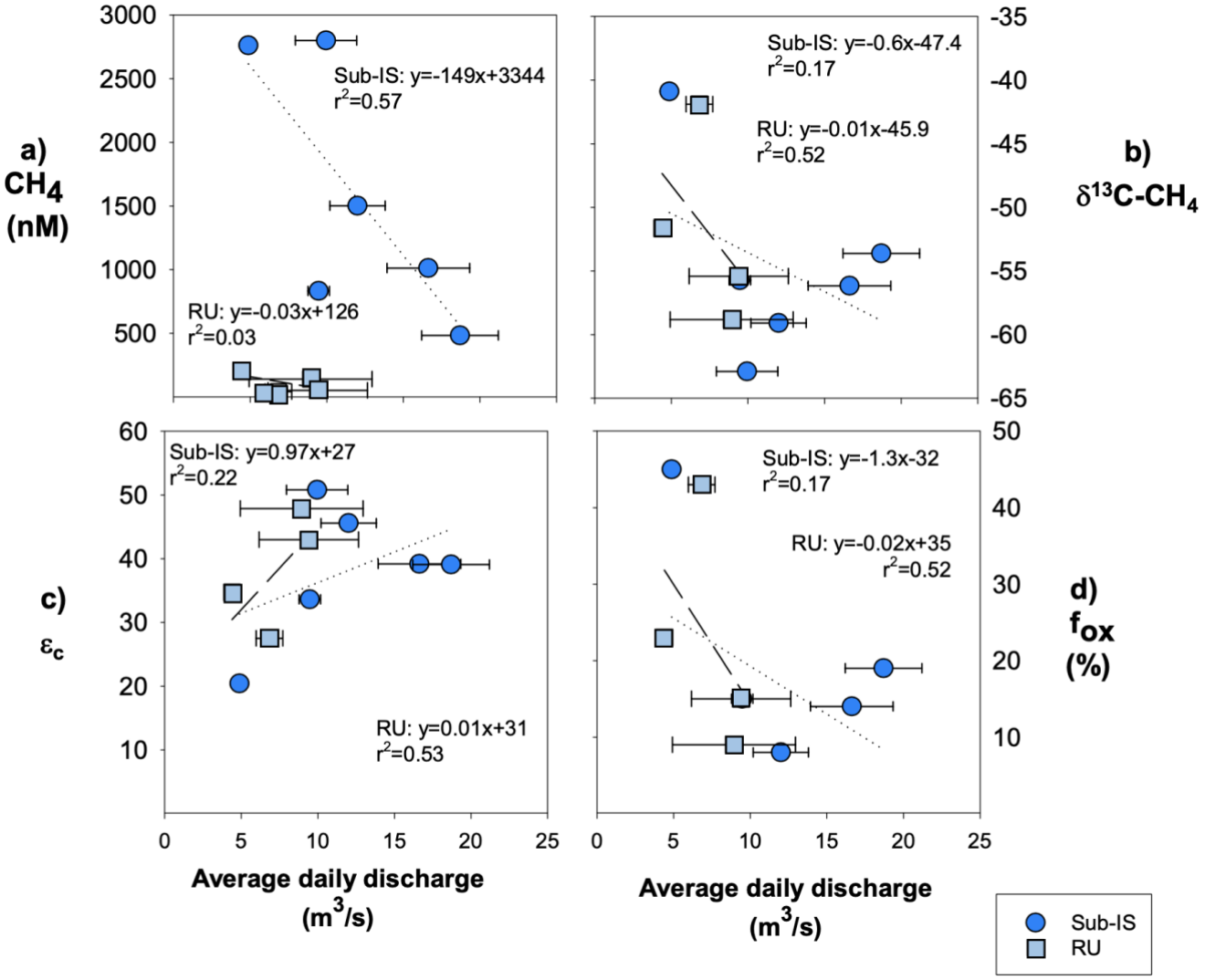

Figure 5. Relationships between average daily discharge and CH_4 dynamics including a) CH_4 concentrations, b) $\delta^{13}C$-CH_4, c) f_{ox}, and d) ε_c for Isunnguata (IS) and Russell (RU) samples. Regressions are shown by dotted lines for Isunnguata and dashed lines for Russell samples. Horizontal error bars represent the standard deviation of average daily discharge for days samples were collected and are smaller than symbols for some data points.

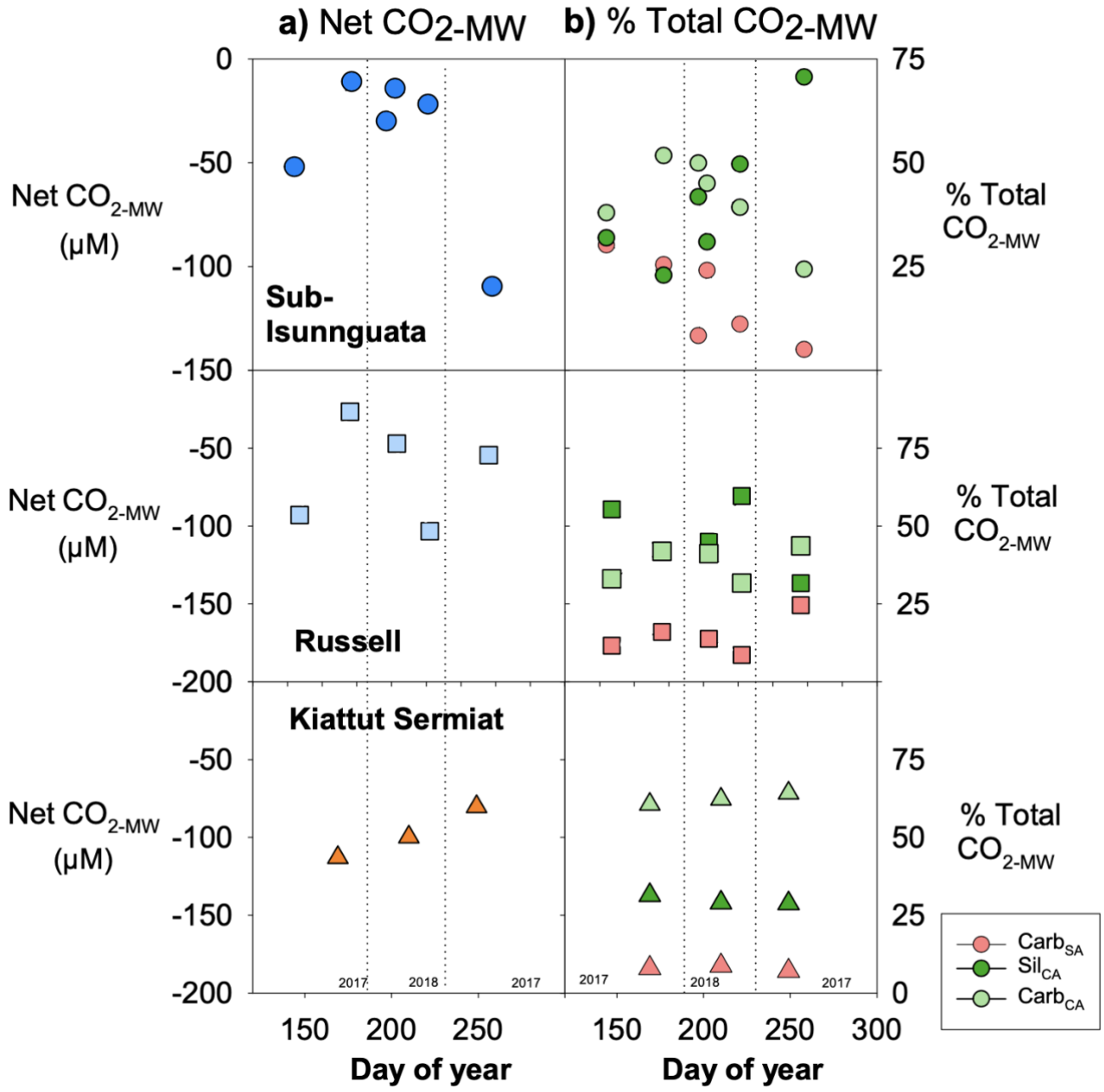

Figure 6. Mineral weathering model results in a) net impact of mineral weathering reactions on CO_2 (Net $CO_{2\text{-MW}}$; Eq. 7) for Isunnguata, Russell, and Kiattut Sermiat subglacial discharge sites (where negative values of Net $CO_{2\text{-MW}}$ indicate net sequestration of CO_2 due to mineral weathering), and b) the proportional contribution of each mineral weathering reaction to the total change in CO_2 (% Total $CO_{2\text{-MW}}$ Eq. 9a-9c).

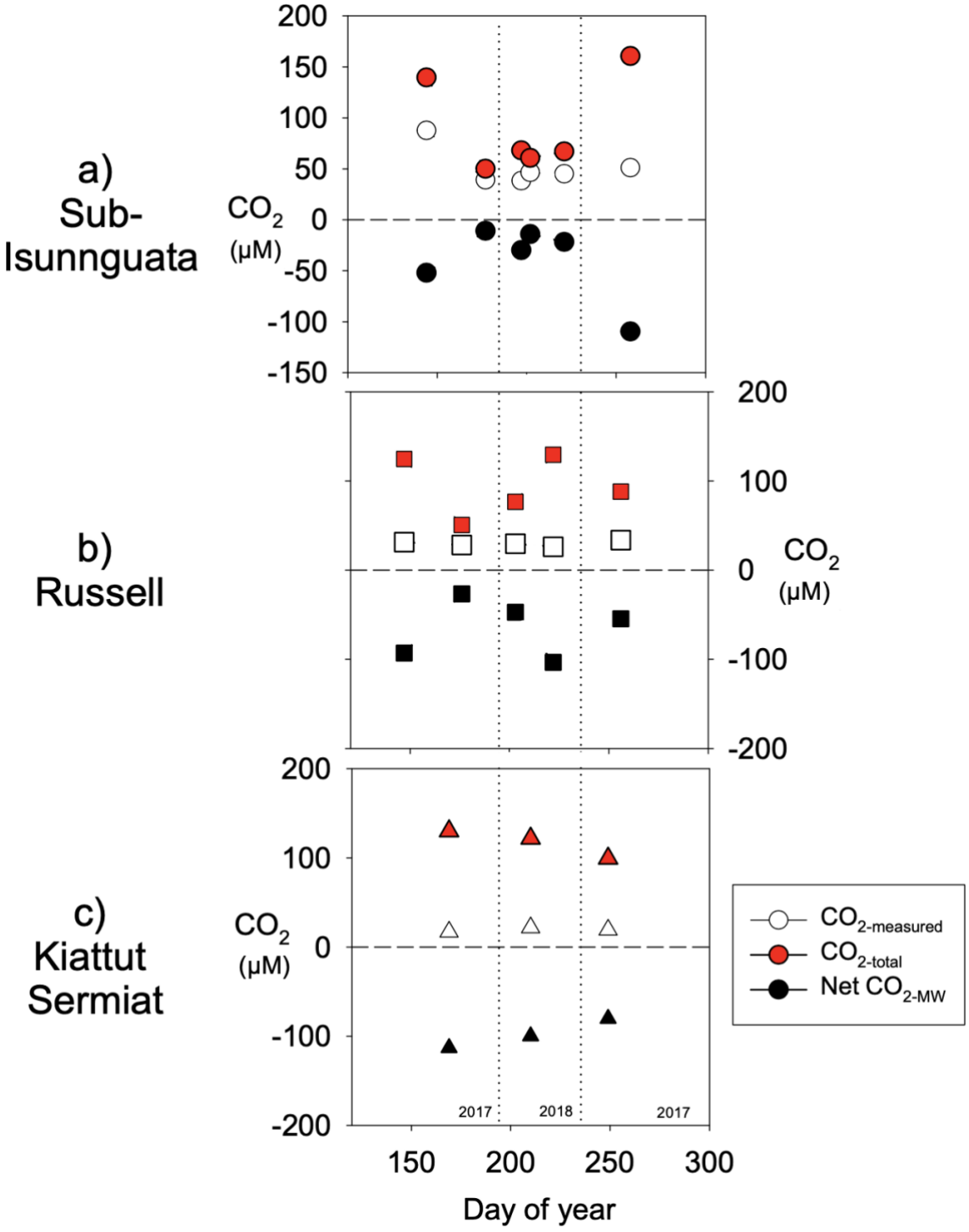

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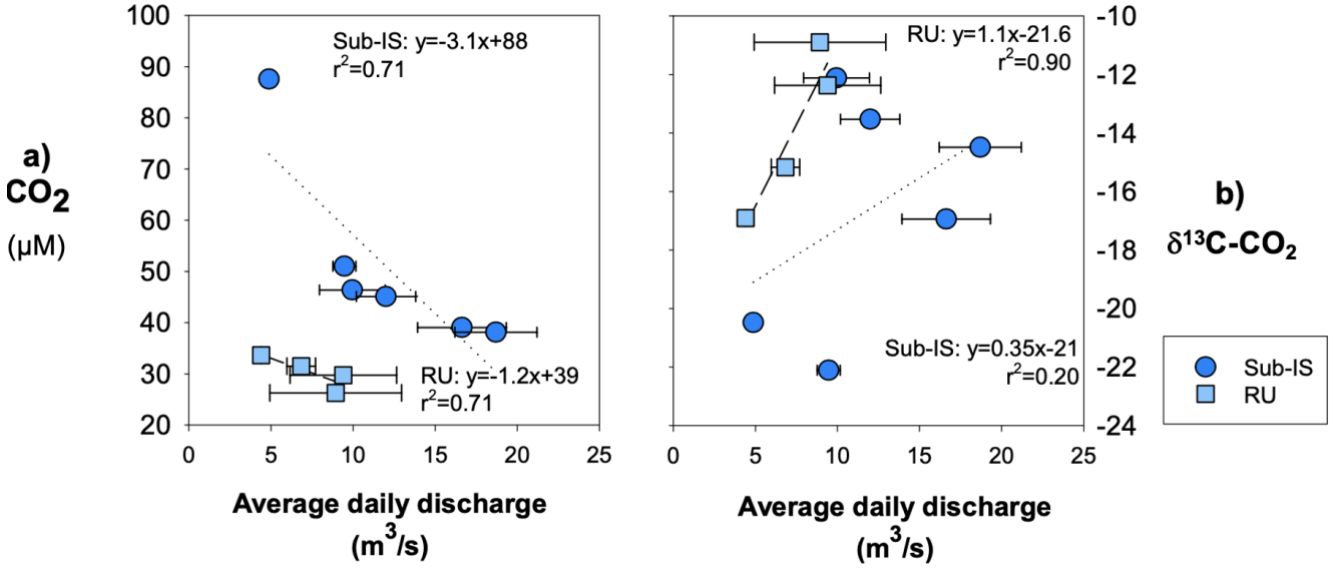
Figure 7. Calculated $CO_{2\text{-total}}$ values for a) sub-Isunnguata, b) Russell, and c) Kiattut Sermiat subglacial discharge against day of the year.



765 **Figure 8. Relationships between average daily discharge a) CO_2 concentrations, and b) $\delta^{13}C\text{-}CO_2$. Regressions are shown by dotted lines for sub-Isunnguata and dashed lines for Russell samples. Horizontal error bars represent the standard deviation of average daily discharge for days samples were collected and are smaller than symbols for some data points.**

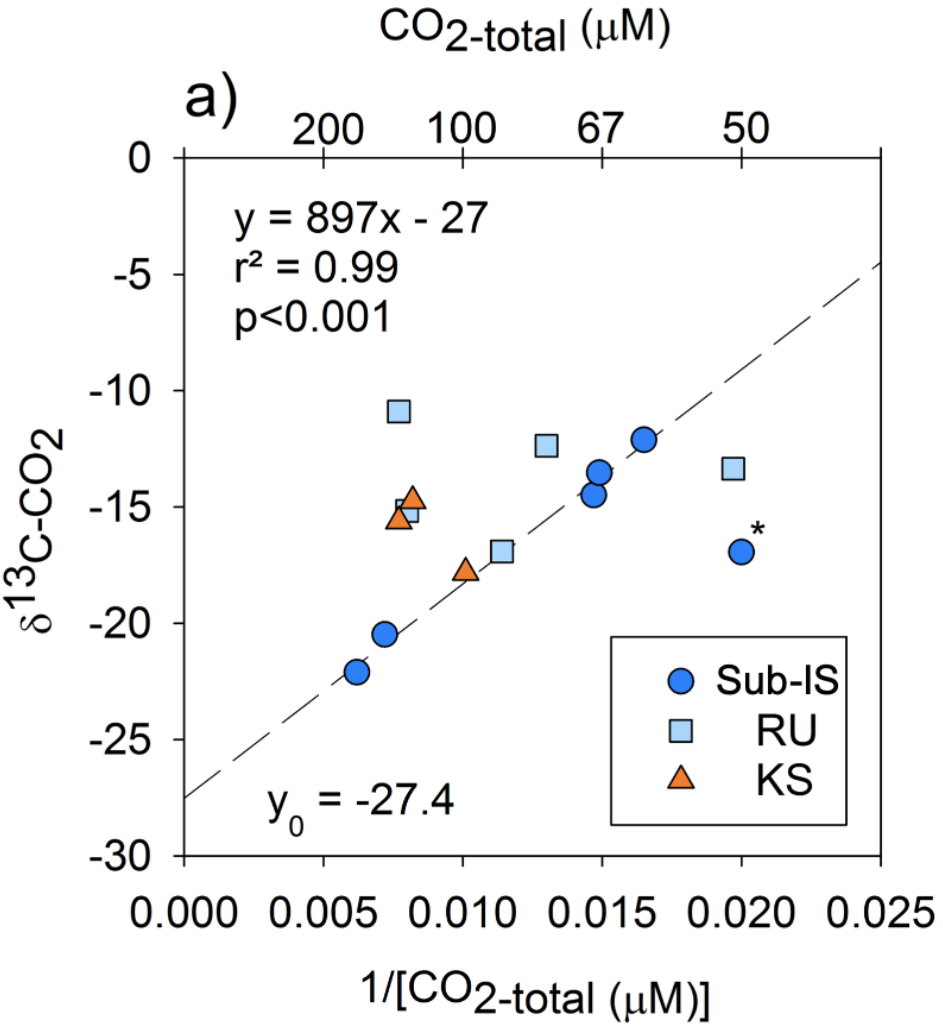


Figure 9. Keeling plot indicating correlations between the magnitude of $CO_{2\text{-total}}$ and $\delta^{13}C\text{-}CO_2$ and for sub-Isunnguata (sub-IS), Russell (RU) and Kiattut Sermiat (KS) samples. Asterisk denotes the outlier not included in the regression between $CO_{2\text{-total}}$ and $\delta^{13}C\text{-}CO_2$ for sub-Isunnguata samples. The plotted regression line was constructed using the sub-IS samples only.